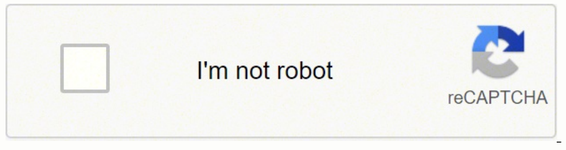

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Uv visible spectroscopy mcqs pdf printable worksheets printable

Answer: Assertion is correct, reason is correct; reason is a correct explanation for assertion. • e.g. UV (185 - 400 nm) / Visible (400 - 800 nm) Spectroscopy, IR Spectroscopy (0.76 - 15 µm) 12. Beer's Law A E .C .l I I T OR log T log A I0 I0 From the equation it is seen that the absorbance which is also called as optical density (OD) of a solution in a container of fixed path length is directly proportional to the concentration of a solution. Chromophore To interpretate UV – visible spectrum following points should be noted: 1. Damkondwar January 21, 2013 2. This is another option for an Overview slide. 34. Then 23. Spectroscopy • It is the branch of science that deals with the study of interaction of matter with light. + - NH2 + NH 3 Cl H A c id ic m e d iu m Aniline λmax = 280 nm λmax = 265 nm 48. If the above compound was completely reduced, what will be the possible transitions it can undergo?(A) n→p* (B) s→s*(C) p→p*(D) n→s*, s→s* 4. Electromagnetic Radiation Violet 400 - 420 nm Yellow 570 - 585 nm Indigo 420 - 440 nm Orange 585 - 620 nm Blue 440 - 490 nm Red 620 - 780 nm Green 490 - 570 nm 9. N N CH3 Pyridine 2-methyl pyridine λmax = 257 nm λmax = 260 nm 49. • Let I be the intensity of incident radiation. Auxochrome e.g. Benzene λmax = 255 nm OH Phenol λmax = 270 nm NH2 Aniline λmax = 280 nm 43. (A) Phenol shows more absorption than phenoxide anion(B) Phenoxide anion shows more absorption than phenol(C) Phenol has equal absorption as phenoxide anion(D) Phenol shows absorption but phenoxide anion doesn't 5. Conjugation of C=C and carbonyl group shifts the λmax of both groups to longer wavelength. C be the concentration of the solution. Sharma • Physical Chemistry – Puri, Sharma & Pathaniya 55. – It can be used to find out molar concentration of the solute under study. R. Then 19. 33. 16. • e.g. Alkenes generally absorb in the region 170 to 205 nm. e.g. NO2, N=O, C=O, C=N, C≡N, C=C, C=S, etc 38. • Determination of molecular weight using Beer's law. Electromagnetic Radiation • Frequency (ν): – It is defined as the number of times electrical field radiation oscillates in one second. On increasing pH, phenol shows (A) Q,R(B) P,R(C) Q,S(D) P,S By increasing the pH to alkaline side, phenol exists as phenoxide ion which shows more absorption due to extra lone pair of electrons participating in pi to pi transition. Therefore it results in bathochromic shift or red shift. • Detection of impurities: – It is one of the important method to detect impurities in organic solvents. What will the audience be able to do after this training is complete? Pavia • Elementary Organic Spectroscopy – Y. 53. Discuss outcomes of the case study or class simulation.Cover best practices. • These transitions usually requires less energy than σ → σ* transitions. Principles of Spectroscopy 10. Principles of Spectroscopy • The principle is based on the measurement of spectrum of a sample containing atoms / molecules. NEET Physics Electromagnetic Spectrum MCQs with answers available in Pdf for free download. Principles of Spectroscopy 1. Select the correct statement regarding their absorption. Term I and Term II. 1 • Bathochromic Shift (Red Shift) • In alkaline medium, p-nitrophenol shows red shift. Add slides to each topic section as necessary, including slides with tables, graphs, and images. 4 • Hypochromic Effect • When absorption intensity (ε) of a compound is decreased, it is known as hypochromic shift. At alkaline pH, phenol loses one proton and exists as phenoxide anion as shown below. Sections can help to organize your slides or facilitate collaboration between multiple authors.NotesUse the Notes section for delivery notes or to provide additional details for the audience. • Spectrum is a graph of intensity of absorbed or emitted radiation by sample verses frequency (ν) or wavelength (λ). Because negatively charged oxygen delocalizes more effectively than the unshared pair of electron. a) Gamma rays b) Ultraviolet rays c) Infrared rays d) Microwaves Answer: Gamma rays Question: The velocity of all radio waves in free space is 3 × 108 m/s. 28. Lambert's Law I 2 . 303 A E .l Lambert's Law 21. • e.g. Mass Spectroscopy 13. See next section for sampletable, graph, image, and video layouts. Run a test print to make sure your colors work when printed in pure black and white and grayscale.Graphics, tables, and graphsKeep it simple: If possible, use consistent, non-distracting styles and colors.Label all graphs and tables. Beer's Law 22. Principle • The UV radiation region extends from 10 nm to 400 nm and the visible radiation region extends from 400 nm to 800 nm. APPLICATIONS OF UV / VISIBLE SPECTROSCOPY 52. When double bonds are conjugated in a compound λmax is shifted to longer wavelength. Give a brief overview of the presentation. 5 • σ → π* transition & • π → σ* transition 6 • These electronic transitions are forbidden transitions & are only theoretically possible. REFERENCES 54. 3 • n → σ* transition • Saturated compounds containing atoms with lone pair of electrons like O, N, S and halogens are capable of n → σ* transition. 3. 303 log K .C .l I I0 K log C .l I 2 . OR • It is the branch of science that deals with the study of interaction of electromagnetic radiation with matter. Absorption Spectroscopy: • An analytical technique which concerns with the measurement of absorption of electromagnetic radiation. • Compounds containing multiple bonds like alkenes, alkynes, carbonyl, nitriles, aromatic compounds, etc undergo π → π* transitions. 2 • Hypsochromic Shift (Blue Shift) • When absorption maxima (λmax) of a compound shifts to shorter wavelength, it is known as hypsochromic shift or blue shift. Interaction of EMR with matter 3. Beer's Law dI C .I dx dI So, K ' C .I dx Integrate equation between limit I = Io at x = 0 and I = I at x=l, We get, I ln K ' C .l I0 24. 4 • n → π* transition • An electron from non-bonding orbital is promoted to anti-bonding π* orbital. x be the thickness of the solution. • The UV spectrum is of only a few broad of absorption. Rotational Energy Levels: • These energy levels are quantized & discrete. 303 Where, log I 0 A Absorbance I K Molar extinction E 2 . a) 10–4 cm b) 10–6 cm c) 10–5 cm d) 10–7 cm Answer: 10–4 cm Question: The wavelength of X-ray is of the order of a) 1 angstrom b) 1 cm c) 1 micron d) 1 metre Answer: 1 angstrom Question: Radio waves do not penetrate in the band of a) ionosphere b) mesosphere c) troposphere d) stratosphere Answer: ionosphere Question: What is the cause of "Green house effect"? Electronic Transitions 29. b) Assertion is correct, reason is correct; reason is not a correct explanation for assertion c) Assertion is correct, reason is incorrect d) Assertion is incorrect, reason is correct. e.g. 1,5 - hexadiene has λmax = 178 nm 2,4 - hexadiene has λmax = 227 nm CH 2 CH 3 H 2C H 3C 40. 42. 1 Hz = 1 cycle per second • Wavelength (λ): – It is the distance between two nearest parts of the wave in the same phase i.e. distance between two nearest crest or troughs. 45. 36. The types of transitions possible in UV-visible region for a compound with molecular formula C 2H 4O are (A) P,Q,R(B) Q,R (C) P,S(D) P,Q,R,S 3. 35. The possible electronic transitions can graphically shown as: 30. • Spectrometer is an instrument design to measure the spectrum of a compound. View these notes in Presentation View during your presentation. 32. • Thus, n → π* & π → π* electronic transitions show absorption in region above 200 nm which is accessible to UV-visible spectrophotometer. • The effect is due to presence of an auxochrome or by the change of solvent. • e.g. Methane (CH4) has C-H bond only and can undergo σ → σ* transition and shows absorbance maxima at 125 nm. Refer to more topic wise NEET Physics Questions and also download more latest study material for all subjects and do free NEET Physics Mock Test Electromagnetic Spectrum NEET Electromagnetic Waves MCQ NEET Electromagnetic Waves students should refer to the following multiple-choice questions with answers for Electromagnetic Spectrum in NEET. Summarize presentation content by restating the important points from the lessons.What do you want the audience to remember when they leave your presentation?Save your presentation to a video for easy distribution (To create a video, click the File tab, and then click Share. Under File Types, click Create a Video.) - - O + O O + O N N - OH A lk a lin e m e d iu m - OH O p-nitrophenol λmax = 255 nm λmax = 265 nm 46. Auxochrome The functional groups attached to a chromophore which modifies the ability of the chromophore to absorb light , altering the wavelength or intensity of absorption. Electromagnetic Radiation 4. Select the wavelength range corresponding to UV-visible region. 11. 2 • π → π* transition • π electron in a bonding orbital is excited to corresponding anti-bonding orbital π*. PRINCIPLES OF UV - VISIBLE SPECTROSCOPY 27. Principles of Spectroscopy 2. Reference Books • Introduction to Spectroscopy – Donald A. The frequency of a radio wave of wavelength 150m is a) 2 MHz b) 1 MHz c) 2 kHz d) 20 kHz Answer: 2 MHz Question: Microwaves are electromagnetic waves with frequency in the range of a) Giga hertz b) Hertz c) Micro hertz d) Mega hertz Answer: Giga hertz Question: Ozone layer above earth's atmosphere will not a) reflect back radio waves b) prevent ultraviolet rays from sun c) prevent infra red radiations originated from earth from escaping earth's atmosphere d) prevent infrared radiations from sun reaching earth Answer: reflect back radio waves Question: The frequency of electromagnetic wave, which best suited to observe a particle of radii 3 × 10–4 cm is of the order of a) 1015 b) 1014 c) 1013 d) 1012 Answer: 1015 Click for more Electromagnetic Waves Study Material › Ministry of Education, Govt of India vide letter No. F.No. 12-5/2020-IS-4 dated 16.12.2021 has intimated that under the banner Azadi ka Amrit Mahotsav the National Yogasanasports Federation has decided to run a project of 750 million Surya Namaskar from 01 January 2022... Emission Spectroscopy: • An analytical technique in which emission (of a particle or radiation) is dispersed according to some property of the emission & the amount of dispersion is measured. Non-conjugated carbonyl group compound give a weak absorption band in the 200 - 300 nm region. 303 log Kl I0 I K log l I0 2 . This results in decrease energy gap between HOMO and LUMO of phenoxide ion. Briefly describe each objective how the audiencewill benefit from this presentation. • The energy required is large for this transition. Add slides to each topic section as necessary, including slides with tables, graphs, and images. 2 • Hypsochromic Shift (Blue Shift) • Aniline shows blue shift in acidic medium, it loses conjugation. Term I... Science is inexplicably linked with our lives and helps us to understand the world around us better. Use a section header for each of the topics, so there is a clear transition to the audience. Electromagnetic Radiation • Electromagnetic radiation consist of discrete packages of energy which are called as photons. 26. Multiple Choice Questions form important part of competitive exams and NEET exam and if practiced properly can help you to get higher rank. a) Ultraviolet rays b) Radio waves c) Infrared rays d) Micro waves Answer: Ultraviolet rays Question: Which one of the following has the shortest wavelength? • The spacing between energy levels are relatively small i.e. 0.01 to 10 kcal/mole. 303 I0 Where, log A Absorbance I K E Absorption coefficient 2 . • Compounds containing double bond involving hetero atoms (C=O, C≡N, N=O) undergo such transitions. 39. 5. • The spacing between energy levels are even smaller than vibrational energy levels. OR The functional groups containing multiple bonds capable of absorbing radiations above 200 nm due to n → π* & π → π* transitions. The MCQ Questions for NEET Physics with answers have been prepared as per the latest NEET Physics syllabus, books and examination pattern. 1. 1 • σ → σ* transition • σ electron from orbital is excited to corresponding anti-bonding orbital σ*. (A) 400-800 nm (B) 200-800 nm (C) 25 µm-2.5 µm(D) 2.5 µm – 1mm 2. • The common solvent used for preparing sample to be analyzed is either ethyl alcohol or hexane. • Detection of isomers are possible. Lambert's Law 18. Keep in mind the font size (important for accessibility, visibility, videotaping, and online production)Coordinated colors Pay particular attention to the graphs, charts, and text boxes.Consider that attendees will print in black and white or grayscale. Electromagnetic Radiation 8. • When the molecules absorb UV-visible light from EMR, one of the outermost bond / lone pair electron is promoted to higher energy state such as E1, E2, ...En, etc is called as electronic transition and the difference is as: ΔE = h ν = En - E0 where (n = 1, 2, 3, ... etc) ΔE = 35 to 71 kcal/mole 15. Question: The ozone layer in the atmosphere absorbs a) X-rays and ultraviolet rays b) only the radiowaves c) only the visible light d) None of these Answer: X-rays and ultraviolet rays Question: Which one of the following has the maximum energy? Interaction of EMR with matter 2. 1. • If auxochrome introduces to the compound, the intensity of absorption increases. 2. Beer's Law I0 2 . UV / VISIBLE SPECTROSCOPY Mr. Santosh M. Non-conjugated alkenes show an intense absorption below 200 nm & are therefore inaccessible to UV spectrophotometer. • n → π* transitions require minimum energy and show absorption at longer wavelength around 300 nm. Shifts and Effects Hyperchromic shift Blue Red Absorbance (A) shift shift Hypochromic shift λmax Wavelength (λ) 51. • The effect is due to presence of an group causes removal of conjugation or by the change of solvent. You're Reading a Free Preview Page 2 is not shown in this preview. CH3 Naphthalene 2-methyl naphthalene ε = 19000 ε = 10250 50. UNESCO has declared 21st February of every year to be celebrated as International Mother Language day to promote dissemination of Mother Language of all, create awareness of linguistic and cultural traditions and diversity across the world and to inspire solidarity... These MCQ questions with answers for NEET Electromagnetic Waves will come in exams and help you to score good marks Electromagnetic Spectrum MCQ Questions with Answers a) Visible light b) Microwaves c) uv-rays d) Radio waves Answer: Visible light Question: The range of wavelength of visible light is a) 4000 Å to 8000 Å b) 15,000 Å to 25,000 Å c) 50 Å to 500 Å d) 9000 Å to 25,000 Å Answer: 4000 Å to 8000 Å Question: Which of the following is the infrared wavelength? Chromophore The part of a molecule responsible for imparting color, are called as chromospheres. Scientific and technological developments contribute to progress and help improve our standards of living. Lambert's Law • When a monochromatic radiation is passed through a solution, the decrease in the intensity of radiation with thickness of the solution is directly proportional to the intensity of the incident light. Interaction of EMR with Matter 14. Electronic Energy Levels: • At room temperature the molecules are in the lowest energy levels E0. Resources • sch/VirtTxtJml/Spectrpy/UV- Vis/spectrum.htm • E2 %80%93visible_spectroscopy • orials/molspec/uvvisab1.htm This template can be used as a starter file for presenting training materials in a group setting.SectionsRight-click on a slide to add sections. As the energy gap decreases, the wavelength of absorption increases as the two parameters are inversely proportional to each other. OR The functional group with non-bonding electrons that does not absorb radiation in near UV region but when attached to a chromophore alters the wavelength & intensity of absorption. Chromophore e.g. O Acetone which has λmax = 279 nmO C H 3C CH3 and that cyclohexane has λmax = 291 nm. – The unit for frequency is Hertz (Hz). a) Infrared rays b) Ultraviolet rays c) Radio waves d) X-rays Answer: Infrared rays Question: The waves which are electromagnetic in nature are a) light waves and X-rays b) sound waves and water waves c) water waves and radio waves d) sound waves and light waves Answer: light waves and X-rays Question: Assertion : The basic difference between various types of electromagnetic waves lies in their wavelength or frequencies.Reason : Electromagnetic waves travel through vacuum with the same speed a) Assertion is correct, reason is correct; reason is a correct explanation for assertion. Datesheet for CBSE Board Exams Class 10 (Scroll down for Class 12 Datesheet) Datesheet for CBSE Board Exams Class 12 Ministry of Education (MoE), Government of India has launched a platform for offering Massive Open Online Courses (MOOCs) that is popularly known as SWAYAM (Study Webs of Active learning for Young Aspiring Minds) on 9 th July, 2017. Give a brief overview of the presentation. This is another option for an Overview slides using transitions. • e.g. when IR radiation is absorbed, molecules are excited from one vibrational level to another or it vibrates with higher amplitude. This is another option for an Overview slide. ΔErotational < ΔEvibrational < ΔEelectronic 17. Electromagnetic Radiation • The relationship between wavelength & frequency can be written as: c=νλ • As photon is subjected to energy, so E=hν=hc/λ 7. • The number of organic functional groups with n → σ* peaks in UV region is small (150 – 250 nm). Applications • Qualitative & Quantitative Analysis: – It is used for characterizing aromatic compounds and conjugated olefins. • e.g. An auxochrome group like –OH, -OCH3 causes absorption of compound at longer wavelength. 6. This decision was taken due to the uncertainty arising out of COVID 19 Pandemic. Absorption & Intensity Shifts 44. • A photon consists of an oscillating electric field (E) & an oscillating magnetic field (M) which are perpendicular to each other. The possible electronic transitions are 1 • σ → σ* transition 2 • π → π* transition 3 • n → σ* transition 4 • n → π* transition 5 • σ → π* transition 6 • π → σ* transition 31. By engaging with this subject, students learn to think, solve... Chromophore 3. Lambert's Law dI I dx dI So, KI dx Integrate equation between limit I = Io at x = 0 and I = I at x=l, We get, I ln Kl I0 20. Interaction of EMR with matter 1. CBSE vide Circular No.Acad-51/2021 dated 5th July, 2021, notified that in the session 2021-2022, Board Examinations would be conducted in two terms, i.e.. Beer's Law • When a monochromatic radiation is passed through a solution, the decrease in the intensity of radiation with thickness of the solution is directly proportional to the intensity of the incident light as well as concentration of the solution. 1 • Bathochromic Shift (Red Shift) • When absorption maxima (λmax) of a compound shifts to longer wavelength, it is known as bathochromic shift or red shift. Vibrational Energy Levels: • These are less energy level than electronic energy levels. Near UV Region: 200 nm to 400 nm Far UV Region: below 200 nm • Far UV spectroscopy is studied under vacuum condition. Terms used in UV / Visible Spectroscopy 37. NCERT now offers online courses for... e.g. Ethylene has λmax = 171 nm O Acetone has λmax = 279 nm C H 2C CH 2 H 3C CH3 Crotonaldehyde has λmax = 290 nm O H 2C C CH3 41. 47. 303 coefficient A E .C .l Beer's Law 25. Describe the major focus of the presentation and why it is important.Introduce each of the major topics.To provide a road map for the audience, you can repeat this Overview slide throughout the presentation, highlighting the particular topic you will discuss next. 3 • Hyperchromic Effect • When absorption intensity (ε) of a compound is increased, it is known as hyperchromic shift.

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